

**activity (relative activity),  $a$**

Defined by the equation  $a = \exp[(\mu - \mu^0)/RT]$  where  $R$  is the gas constant,  $T$  the thermodynamic temperature,  $\mu$  the *chemical potential* and  $\mu^0$  the standard chemical potential the definition of which depends on the choice of the *standard state*.

See also *absolute activity*.

G.B 49; 1996, 68, 989