

isoelectronic

Two or more *molecular entities* are described as isoelectronic if they have the same number of valence electrons and the same structure, i.e. number and *connectivity* of atoms, but differ in some of the elements involved.

Thus:

CO, N₂ and NO⁺ are isoelectronic.

CH₂=C=O and CH₂=N=N are isoelectronic.

CH₃COCH₃ and CH₃N=NCH₃ have the same number of electrons, but have different structures, hence they are not described as isoelectronic.

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