

partial anodic (cathodic) current

When a single electrochemical reaction occurs at an electrode (i.e. all other electrochemical reactions may be neglected):

$$I = I_a + I_c$$

where I is the total current, I_a is the positive partial anodic current and I_c is the negative partial cathodic current. When more than one reaction is significant the reactions may be numbered and the numbers used as subscripts: $I_{1,a}$, $I_{1,c}$, $I_{2,a}$, etc. Then:

$$I = \sum I_{i,a} + \sum I_{i,c}$$

1974, 37, 512